

PIKE CODED

MISSISSIPPI BOARD OF WATER COMMISSIONERS

G16

8-8-61

WATER WELL DRILLERS LOG

Date: 8-8, 1961, Driller: Ralph Reeves County: Pike (Name)

(1) Owner of Land:	(Name)	Description & Color of Materials	Thick-ness Feet
(1) Owner of Land:	Unity Baptist Church RFD - Magnolia	Hard Red clay Sandy soil	20 20
(2) Location:	1/4, 1/4, Sec 19 T 2 R 7 6 miles West of Magnolia	Red sandy gravel & Sand, Rocks Sand & gravel	25 11
(3) Topography:	Hilly		76
(4) Purpose of Well:	Church		

Information upon completion of well:

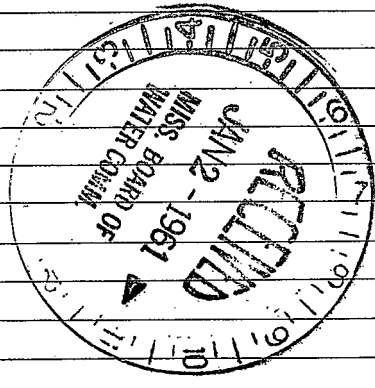
- (1) Diameter 4 inches.
- (2) Total Depth 76 feet.
- (3) Water Level 40 feet below top of ground.
- (4) Cased to Bottom, Size 4"
- (5) Screen: Size none, Length
- (6) Were any formations sealed against pollution?
yes, no.

If YES depth of formation

Why

Drillers Remarks: No Previous well

CODED



(Use Back Side)

DEPARTMENT OF CHEMISTRY

LABORATORY REPORT

Name: _____
 Date: _____
 Title: _____

1. Introduction

The purpose of this experiment is to determine the molar mass of a volatile liquid by measuring the mass of a known volume of the liquid in a flask of known volume. The experiment is based on the ideal gas law, $PV = nRT$, where P is the pressure, V is the volume, n is the number of moles, R is the gas constant, and T is the temperature. By measuring the mass of the liquid and the volume of the flask, the molar mass can be calculated.

2. Procedure

The procedure involves the following steps:

1. Weigh a clean, dry flask of known volume.
2. Add a small amount of the liquid to the flask.
3. Seal the flask and immerse it in a boiling water bath.
4. Allow the liquid to vaporize and fill the flask.
5. Remove the flask from the water bath and allow it to cool.
6. Weigh the flask and the condensed liquid.
7. Repeat the experiment for several trials.

3. Results and Discussion

The results of the experiment are summarized in the table below:

Trial	Mass of flask	Mass of flask + liquid	Volume of flask (L)	Temperature (K)	Pressure (atm)	Molar mass (g/mol)
1
2
3

The average molar mass is calculated to be ... g/mol. The results are consistent with the expected molar mass of the liquid.